

WORKSHEET2

SUBJECT-CHEMISTRY

- List five physical properties of metals and compare them with non-metals.
- Why is gold widely used for making jewellery?
- Name one metal commonly used for making cooking utensils. Give reason also.
- Give exceptions in the following cases-
 - All metals exist as solid at room temperature.
 - Non-metals are non-lustrous.
 - Non-metals do not conduct electricity.
 - Metals are hard.
 - Solid non-metals are brittle.
- What is observed when-
 - Magnesium ribbon is burnt in a flame.
 - Copper metal is heated in air.
- Name two metal oxides that are soluble in water. (Write equation also)
- With the help of equations, show that Al_2O_3 is an amphoteric oxide.
- Write equations for the reactions of an acid with:
 - Zn metal.
 - Na_2CO_3
 - $NaHCO_3$
 - $NaOH$ solution.
- Why is there no evolution of hydrogen when nitric oxide reacts with metals?
- What is the reactivity series of metals?
- What is an electrovalent bond? Show the formation of $NaCl$, Na_2O , MgO and $MgCl_2$ with the help of electron-dot representation.
- Why does a solution of sodium chloride conduct electricity which solid $NaCl$ does not?
- Write properties of ionic compounds.
- Differentiate between a mineral and an ore.
- Give reasons:
 - Gold and silver are found in their free state.
 - Sodium is never found in its free state.
 - The sulphide ore is converted into an oxide for the extraction of metals.
 - The oxides of metals like Hg can be reduced by heating only.
 - The oxides of metals like Na, Mg, Ca cannot be reduced by carbon
- Name two metals that can be refined by electrolysis.
- Name the cathode, anode and the electrolyte for refining of copper electrolytically.
- Name the substance formed on the surface of iron, silver and copper due to corrosion.
- How is steel different from stainless steel?
- What is an alloy? Write the composition of brass, bronze and solder.